

GENERAL CHEMISTRY I

1	Course Title:	GENERAL CHEMISTRY I	
2	Course Code:	KIM1031	
3	Type of Course:	Compulsory	
4	Level of Course:	First Cycle	
5	Year of Study:	1	
6	Semester:	1	
7	ECTS Credits Allocated:	5.00	
8	Theoretical (hour/week):	2.00	
9	Practice (hour/week):	0.00	
10	Laboratory (hour/week):	0	
11	Prerequisites:	None	
12	Language:	Turkish	
13	Mode of Delivery:	Face to face	
14	Course Coordinator:	Prof. Dr. RAHMIYE AYDIN	
15	Course Lecturers:	Yok	
16	Contact information of the Course Coordinator:	rahmiye@uludag.edu.tr Tel: 0 (224) 2941729 Uludağ Üniversitesi Fen-Edebiyat Fakültesi Kimya Bölümü, 16059, BURSA	
17	Website:		
18	Objective of the Course:	The purpose of this course is to teach the structure and properties of matter, atom which the smallest building blocks of matter, the names of the formulas of compounds, chemical reactions and equations, properties and types of reaction occurring in aqueous solution, the properties of gases, gas laws and thermochemistry.	
19	Contribution of the Course to Professional Development:		
20	Learning Outcomes:		
		1	Learn the most basic terms of chemistry, the methods applied while doing chemical measurement and speaking of their results and can apply laboratories studies.
		2	Describe theories of the fundamental laws of chemistry and atomic structure.
		3	Learn to identify the characteristics and behavior of states of matter and the structure, names and formulas of compounds
		4	Learn the properties and the stoichiometry of chemical reactions.
		5	Chemical processes occurring in biological environments, learn to interpret by the basic laws of chemistry.
		6	Investigate the developments in the field of chemistry and transfer in the field of biology.
		7	Apply the knowledge of basic chemistry in the biology and chemistry laboratory.
		8	
		9	
		10	
21	Course Content:		
		Course Content:	

Week	Theoretical	Practice		
1	PROPERTIES AND MEASUREMENT OF MATTER: The purpose of chemistry, scientific method, properties and classification of matter,Measurement of matter, uncertainty in the scientific method, significant figures.			
2	ATOMS AND ATOMIC THEORY: the first discoveries in chemistry and atomic theory, electrons and other discoveries in atomic physics, atomic nucleus, the chemical elements, atomic masses, periodic table entry.			
3	ATOMS AND ATOMIC THEORY: The mole concept and avogadro number, using the mole concept of calculations.			
4	CHEMICAL COMPOUNDS: Chemical compounds and their formulas, the mole concept and chemical compounds, composition of chemical compounds.			
5	CHEMICAL COMPOUNDS: Oxidation steps, naming chemical compounds, nomenclature and formulas of inorganic and organic compounds.			
6	CHEMICAL REACTIONS: Chemical reactions and chemical equations, chemical equations and stoichiometry.			
7	CHEMICAL REACTIONS: Chemical reactions in solution, determine the limiting component			
Activites		Number	Duration (hour)	Total Work Load (hour)
9	REACTIONS IN AQUEOUS SOLUTIONS: The nature of aqueous solutions, precipitation	14	2.00	28.00
Practicals/Labs		0	0.00	0.00
Self study and preparation		14	2.00	28.00
Homeworks		0	0.00	0.00
Projects		0	0.00	0.00
Field Studies		0	0.00	0.00
Midterm exams		1	27.00	27.00
11	GASES: Properties of gases: the gas equation and its applications, the gases of chemical reactions, gas mixtures	1	40.00	40.00
Others		0	0.00	0.00
Total Work Load				177.00
Total work load of gas properties related to this course				5.00
ECTS Credit of the Course				5.00
13	Thermochemistry			
14	Thermochemistry			
22	Textbooks, References and/or Other Materials:	General Chemistry I, Petrucci Harwood Herring. Palme Publishing Lecturer course notes		
23	Assesment			
TERM LEARNING ACTIVITIES		NUMBE R	WEIGHT	
Midterm Exam		1	40.00	
Quiz		0	0.00	
Home work-project		0	0.00	
Final Exam		1	60.00	

Total	2	100.00
Contribution of Term (Year) Learning Activities to Success Grade	40.00	
Contribution of Final Exam to Success Grade	60.00	
Total	100.00	
Measurement and Evaluation Techniques Used in the Course		
24	ECTS / WORK LOAD TABLE	

25	CONTRIBUTION OF LEARNING OUTCOMES TO PROGRAMME QUALIFICATIONS															
	PQ1	PQ2	PQ3	PQ4	PQ5	PQ6	PQ7	PQ8	PQ9	PQ10	PQ11	PQ12	PQ13	PQ14	PQ15	PQ16
ÖK1	3	1	3	4	1	3	5	1	3	4	4	5	3	3	4	2
ÖK2	3	1	3	4	1	4	3	1	3	4	3	4	3	3	4	2
ÖK3	5	1	3	5	1	4	4	1	3	4	4	4	3	2	4	3
ÖK4	3	1	3	4	1	5	5	1	3	4	4	4	3	2	3	3
ÖK5	4	1	3	5	1	5	5	1	3	4	4	4	2	2	3	4
ÖK6	4	3	3	5	1	4	4	1	3	4	5	5	3	3	2	3
ÖK7	4	2	3	5	1	3	5	1	3	4	5	5	4	4	3	3
LO: Learning Objectives PQ: Program Qualifications																
Contribution Level:	1 very low		2 low		3 Medium		4 High		5 Very High							