ORGANIC CHEMITRY									
1	Course Title:	ORGAN	IC CHEMITRY						
2	Course Code:	KIM1080							
3	Type of Course:	Compuls	sory						
4	Level of Course:	First Cyc	le						
5	Year of Study:	1							
6	Semester:	2							
7	ECTS Credits Allocated:	3.00							
8	Theoretical (hour/week):	3.00							
9	Practice (hour/week):	0.00							
10	Laboratory (hour/week):	0							
11	Prerequisites:	None	None						
12	Language:	Turkish	Turkish						
13	Mode of Delivery:	Face to face							
14	Course Coordinator:	Prof. Dr. AYHAN YILDIRIM							
15	Course Lecturers:	Prof. Dr. Necdet Coşkun Doç. Dr. Mustafa Tavaslı Doç. Dr. Nevin Arıkan Ölmez							
16	Contact information of the Course Coordinator:	Prof. Dr. Ayhan YILDIRIM yildirim@uludag.edu.tr 0 224 29 41 771 Uludağ Üniversitesi Fen-Edebiyat Fakültesi Kimya Bölümü, Görükle/BURSA 16059							
17	Website:								
18	Objective of the Course:	The aim of the course is to teach the biologically important organic molecules contain functional groups, structural frame and tri- dimensional structural of organic molecules.							
19	Contribution of the Course to Professional Development:	At the end of this course, the student will gain knowledge and skills about professional development and basic organic chemistry as well as; He can be a qualified engineer or researcher who has sufficient knowledge of organic chemistry that our country needs in the field of food engineering, who has the power to solve different problems in the institution where he works.							
20	Learning Outcomes:								
	•	1	Knowing of the Organic Chemistry and being aware of the importance in Biology.						
		2	Learning of the basic Organic Chemistry terms.						
		3	Learning the reactions of some basic Organic functional groups.						
		4	Learning the physical and chemical properties of some organic compounds that being Biological important.						
		5	Learning the bio-effectiveness and/or the bio-harmfulness of some organic compounds and using these with this consciousness						
		6	Being knowledgeable about the applications of some organic compounds existed in nature.						
		7							
		8							
		9							
		10							
04	Course Content:								

	Course Content:											
Week	Theoretical	Practice										
1	Alkanes: - Molecular geometry and Bond angle - Closed, Opened, Compressed and Lineer Formula - Straight chained Alkanes (MethaneDecan)											
2	Alkyl Groups: -General Formula -Methyl, ethyl, n-propyl, n-butyl, n-pentyl, n- hexyln-decyl -İsopropyl, sec-butyl, ter-butyl, sec-pentyl, neopentyl											
3	Branched alkanes: -Structural isomerism -Systematic nomenclature											
4	Alcohols and Ethers: -Water molecule -Molecular geometry and Bond angle Alcohols: -Closed, Opened, Compressed and Lineer Formula -Primary, secondary and tertiary alcohols -Mono and poly alcohols, Systematic nomenclature -Forces that holding the molecules together (H-Bond)											
Activit	es	Number	Duration (hour)	Total Work Load (hour)								
Theore	tEnrces that holding the molecules together	14	3.00	42.00								
Practica	als/Labs	0	0.00	0.00								
Self stu	dynamobrapeodecule	1	14.00	14.00								
Homew	vorks	0	0.00	0.00								
Project	Formula	0	0.00	0.00								
Field S	tudies	0	0.00	0.00								
Midtern	TStrangterral isomerism	1	14.00	14.00								
Others		0	0.00	0.00								
Final E	Charge stat holding the molecules together	1	20.00	20.00								
Total W	/ork Load			90.00								
Total w	ocanadyaghoup			3.00								
ECTS	Credit of the Course			3.00								
	-Closed, Opened, Compressed and Lineer Formula -Systematic nomenclature Ketones: -Closed, Opened, Compressed and Lineer Formula -Systematic nomenclature -Structural isomerism -Forces that holding the molecules together (Dipole-dipole interaction) -Aldose and Ketose sugars											

7	Carboxylic acids and Esters:		
	-Carbonyl, Hydroxyl and Alcoxy grou	ps	
	Carboxylic acids:	•	
	-Closed, Opened, Compressed and L	_ineer	
	Formula		
	-Systematic nomenciature		
	-Forces that holding the molecules to	aether	
	(H-Bond)	3	
	Esters:		
	-Closed, Opened, Compressed and L	₋ineer	
	Formula Systematic nomonclature		
	-Systematic nomenciature	aether	
	(Dipole-dipole interaction)	gettier	
	-Vegetable and animal oils and Waxs	6	
8	Amides:		
	-Carbonyl, Hydroxyl and Amide group	ps	
	-Molecular geometry and Bond angle	incor	
	Formula		
	-Systematic nomenclature		
	-Forces that holding the molecules to	gether	
	(H-Bond or Dipole-dipole interaction)	l .	
	-Aminoacids, Proteins and Peptide b	onas	
9	Alkenes:		
	-Closed, Opened, Compressed and L	_ ineer	
	Formula		
	-Systematic nomenclature		
	-Geometrical isomerism (Cis-Trans		
	-Forces that holding the molecules to	aether	
	(Van der Waals interaction)	geniei	
	-Vegetable fatty acids		
10	Aromatic compounds:		
	-Aromaticity		
	-Benzenoid Aromatic compounds		
	-Ortho-/meta-/ para- positions		
11	Functional group converisons:		
	-Determination of oxidation step		
	-Electronegativity (C, H, O, Cl, Br)		
12	Reduction Reaction:		
40	-Reductive reagents		
13	-Oxidative reagents		
14	Acid-base reaction:		
	-Description of acid and base		
	-Poor acids		
	-Buffer solutions		
22	Textbooks, References and/or Other		i) G. Solomons ve C. Fryhle (Çev. Ed. G. Okay ve Y.
	iviaterials:		i) R. J. Fessenden ve J. S. Fessenden (Cov. Ed. T. Liver)
			Organik Kimva, Günes Kitabevi 1992
			iii) J. McMurry, Organic Chemistry, Brooks/Cole Publishing
			Comp., 1992.
			iv) P. Y. Bruice, Organic Chemistry, Prentice Hall, 2001.
23	Assesment		
TERMI		NUMBE	WEIGHT
R			
Midterr	n Exam	1	40.00

Quiz					0)	0.0	0.00								
Home work-project						0		0.0	0.00							
Final Exam 1							60	60.00								
Total						2		10	0.00							
Contribution of Term (Year) Learning Activities to Success Grade							40	40.00								
Contribution of Final Exam to Success Grade							60	60.00								
Total							10	100.00								
Measurement and Evaluation Techniques Used in th Course						d in th	ne Sh Ora Ho	, Short-term quizzes are held Oral exams are held in the course Homework given								
24 EC	;TS/	WO	RK L	OAD	TAB	LE										
25	25 CONTRIBUTION OF LEARNING OUTCOMES TO PROGRAMME QUALIFICATIONS															
	PQ1	PQ2	PQ3	PQ4	PQ5	PQ6	PQ7	PQ8	PQ9	PQ1 0	PQ11	PQ12	PQ1 3	PQ14	PQ15	PQ16
ÖK1	5	0	0	5	0	4	0	0	0	0	0	0	0	0	0	0
ÖK2	0	0	0	5	0	0	0	0	0	0	0	0	0	0	0	0
ÖK3	0	0	0	5	0	0	0	0	0	0	0	0	0	0	0	0
ÖK4	0	0	0	5	0	0	0	0	0	0	0	0	0	0	0	0
ÖK5	0	0	0	0	0	5	0	0	0	0	0	0	0	0	0	0
ÖK6	0	0	0	5	0	0	0	0	0	0	0	0	0	0	0	0
LO: Learning Objectives PQ: Program Qualifications																
Contrib ution Level:1 very low very low2 low			3 Medium		4 High			5 Very High								